



Palestine Technical University - Kadoori

Faculty of Applied Sciences

Wednesday, 5/12/2016

General Chemistry Lab1

Quiz

Time allowed: 10 Min.

Name (in Arabic):

Section (day & time):

(Molar mass, $\text{BaCl}_2 \cdot 2\text{H}_2\text{O} = 244$, $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O} = 380$, $\text{Ba}_3(\text{PO}_4)_2 = 601$,

(Q1) A 4.83 g sample mixture of $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ and $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$ was boiled in 150 mL distilled water. The precipitate, $\text{Ba}_3(\text{PO}_4)_2$, was separated by filtration, dried and weighed a mass of 1.68 g. Barium ions (Ba^{+2}) were added to filtrate and white precipitate was formed. Answer the following questions:

1- The limiting reactant is.....

2- Write the net ionic equation

3- Mass of the limiting reactant

4- Mass of the excess reactant